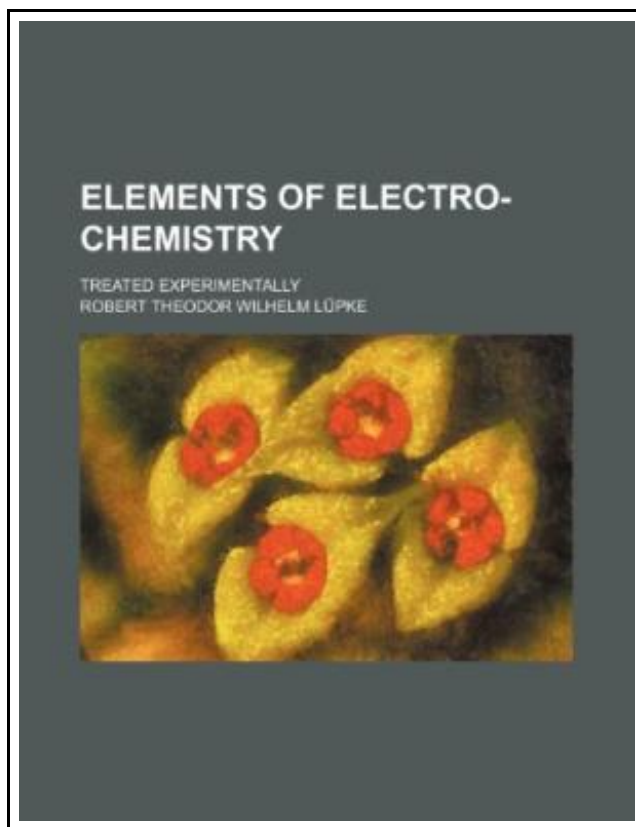


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1903 Excerpt: .can be dependent only on the constants of the solvent. Van t Hoff (see Nernst s Theoretical Chemistry English Ed., pp. 126, 128) elucidated the relations between osmotic pressure and the three other magnitudes; and he established the following equations, in which M is the molecular weight, and s is the specific gravity, of the solvent, T₀ and T₀ are its absolute boiling point and freezing point respectively, w and w are the latent heat of evaporation and the latent heat of fusion of the solvent per gram, and K, K , and K are collective constants depending on the solvents used. I. $P = P_{jza} \times \frac{1}{M} \times \frac{1}{K}$ atmos. Px M Px II. $P = \frac{1}{K} \times \frac{W-s-w}{T_0}$ K atmos. III. $P = \frac{1}{K} \times \frac{1}{l} \times \frac{1}{l} = \frac{1}{K} \times \frac{1}{l} = \frac{1}{K} \times \frac{1}{l}$ K atmos. Hence it must be possible to find one magnitude from the others. Values for P may be found by applying equations I. and II. to the experimental data which have been tabulated already. The results of doing this are to be seen in the two following tables: --In both tables P (in column 11) is calculated from the equation $PV = 0.819 T_0$, or $= 0.819 T_0$, litre-atmospheres; where V is the number of litres which would contain one gram-molecule of the substance in solution. Considering that the solutions which were used in determining boiling points and freezing points were, relatively, very concentrated, it is seen that the values thus calculated for P...



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